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Quantitative calculations involving reactions in solution are

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carried out in the same manner as we discussed in Chapter 11. Instead of masses, however, we use volumes of solutions of known concentration to determine the number of moles of reactants. Whether we are dealing with volumes of solutions of reactants or masses of reactants, the coefficients in the balanced chemical equation tell us ...

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378 Chapter 12 12CHAPTER Study Guide Key Concepts 12.1 The Arithmetic of Equations • A balanced chemical equation provides the ... Stoichiometry 379 CHAPTER 12 Assessment 36. a. Two formula units  $\text{KClO}_3$  decompose to form two formula units  $\text{KCl}$  and three molecules  $\text{O}_2$ . b.

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reactions dependent upon? 3. If I have ten hydrogen molecules and three oxygen molecules, how many molecules of water can I make?

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and effect The runner started the race with energy and excitement, but as she crossed the

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Stoichiometry CHAPTER 12 What You'll Learn You will write mole ratios from balanced chemical equations. You will calculate the number of moles and the mass of a reactant or product when given the number of moles or the mass of another reactant or product. You will identify the limiting reactant in a chemical reaction. You will determine the

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