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264 CHAPTER 8. A coefficient multiplies the number of atoms of each element indicated in a chemical formula. Thus,  $2\text{H}_2\text{O}$  represents four H atoms and two O atoms. To add two more hydrogen atoms to the right side of the equation, one may be tempted to change the subscript in the formula of water so that H

### **CHAPTER 8 Chemical Equations and Reactions**

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Chapter 11 Review Sheet Date \_\_\_\_ Be sure to show all your work for each problem. 1. Calcium carbonate reacts with 14.8g of hydrochloric acid. How many moles of calcium chloride are produced? RXN:  $\text{CaCO}_3 + 2\text{HCl} \rightarrow \text{CaCl}_2 + \text{H}_2\text{CO}_3$ . 14.8 g HCl 1 mol HCl 1 mol  $\text{CaCl}_2 = 0.2\text{ mol CaCl}_2$ . 36.5 g HCl 2 mol HCl. 2.

### CHAPTER 11: STOICHIOMETRY

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Chapter 08 Review Worksheet 2. Empirical Formula Worksheet. Chapter 09 Stoichiometry Hints for solving Stoichiometry Problems . This document was developed by S.E. Van Bramer for Chemistry 145 at Widener University. Chapter 10 Electrons and the Periodic Table

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CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N<sub>2</sub> are mixed with 12.0 mol of H<sub>2</sub> according to the following equation: N<sub>2</sub>(g) + 3H<sub>2</sub>(g) ...

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SCH 3UI - Chapter 5 (Stoichiometry) Review Answer Sheet KNOWLEDGE AND UNDERSTANDING A. Completion. Complete each sentence or statement. 1. A(n) \_\_\_ excess \_\_\_ reagent is the only reagent left over when all other reagents have been consumed in a chemical reaction. 2. A(n) \_\_\_ limiting \_\_\_ is a reagent that is completely consumed in a reaction and thereby limits the amount of product(s) that ...

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