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Chapter 4 Notes - Types of Chemical

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Reactions and Solution Chemistry 4.1  
Water, the Common Solvent A. Structure  
of water 1. Oxygen's electronegativity is  
high (3.5) and hydrogen's is low (2.1) 2.

## **Chapter 4 Notes - Types of Chemical Reactions and Solution ...**

AP Chem Chapter 4. General Chemistry  
ebbings-gammon chapter 4 vocabulary.

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STUDY. PLAY. electrolyte. a substance that dissolves in water to give a solution that conducts electric current.

nonelectrolyte. a substance that dissolves in water to give a nonconducting or very poorly conducting solution.

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AP Chem chapter 4. STUDY. PLAY.  
aqueous solution. a solution in which water is the solvent. solution. a mixture of substances that has a uniform composition; a homogenous mixture. solvent. the dissolving medium of a solution; it is normally the component of



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a solution present in the greater amount.

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AP Chemistry Chapter 4 Oxidation  
number of Hydrogen (H) Oxidation  
number of Oxygen (O) Oxidation number  
of Fluorine (F)

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## **CHAPTER 4 ANSWERS AND EXPLANATIONS - Bonding and Phases ...**

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Identify the choice that best completes  
the statement or answers the question.

\_\_\_ 1. What mass of silver nitrate,  $\text{AgNO}_3$ , is required to prepare 800. g of 3.50%

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solution of  $\text{AgNO}_3$ ? a. 24.6 g b. 26.7 g c.  
27.0 g d. 25.5 g e. 28.0 g \_\_\_\_ 2.

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Objectives. Explain why water is polar  
and how it dissolves ionic compounds  
(4:32) Explain the difference between  
electrolytes and non-electrolytes ...

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## **Chapter 4 - AP Chemistry**

AP Chemistry Practice Test: Chapter 4

Name \_\_\_\_\_

SHORT ANSWER. Write the word or phrase that best completes each statement or answers the question.

1) What is the solvent in an aqueous solution?

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## **AP Chemistry Practice Test: Chapter 4 SHORT ANSWER. Write ...**

AP CHEMISTRY CHAPTER 4 REVIEW (100 points) (6 points) 1. Determine whether the following compounds are soluble or insoluble in water: a.  $\text{Hg}_2\text{Cl}_2$ . b.  $\text{KI}$ . c.  $\text{Pb}(\text{NO}_3)_2$ . d.  $\text{NaBr}$ . e.  $\text{Ba}(\text{OH})_2$ . f.  $\text{CrCO}_3$  (4 points) 2. Determine the precipitate

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that will form, if any, when the following solutions are mixed:

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Configuration 6 - Periodicity. 7 -  
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Chapter 4 Reactions in Aqueous Solution

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(Sections 4.1 - 4.4)

## **Chapter 4 Reactions in Aqueous Solution (Sections 4.1 - 4.4)**

AP Chem Chapter 7,8,9 test ; AP Chem  
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Test Chapter 1, 2, and 3.1

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